

# AFR-related publications

*Complete list of published scientific papers that utilize  
Corning® Advanced-Flow™ Reactor Technology*

Publication Date: March 2022

This list compiles scientific papers published that utilize Corning® Advanced-Flow™ Reactor (AFR) Technology. Corning expresses our thanks to all authors who used our technology in their experiments. If you would like to include your published paper using AFR Technology, please contact us at [reactors@corning.com](mailto:reactors@corning.com) and we will be pleased to review your submission for inclusion in this document.

## 1. Reactors: Goal, Design & Characterization

Corning developed AFR technology to support the synthetic industry as an ongoing effort toward process intensification.<sup>1</sup> For this, switching the synthetic paradigm<sup>2</sup> from traditional batch to flow chemistry was pursued.<sup>3</sup> The reactors were designed towards high scale production<sup>4-10</sup> with a variety of applications.<sup>11</sup>

## 2. Reactors Engineering & Characterization

Using Corning's expertise, reactors were designed either in resistant glass<sup>12</sup> or Silicon Carbide (no chemical limitation found yet).<sup>13</sup> The mass transfer properties,<sup>14</sup> heat exchange,<sup>15</sup> pressure drop<sup>16</sup> and residence time distribution<sup>17</sup> were fully characterized for single<sup>18,19</sup> or dual phase systems.<sup>20-22</sup> The hydrodynamic properties of liquid and gas liquid<sup>23</sup> flow were published.<sup>24,25</sup> The same work was also carried out for the Corning® Low-Flow Reactor.<sup>26</sup> Light was also characterized in photochemical reactors.<sup>27</sup>

To help with industrialization, the design of Corning's reactors ensured a scalable system such as liquid/liquid systems from Low-Flow to G1<sup>28</sup>, and up to production.<sup>29,30</sup> The concept behind flow reactors and scale-up has been summarized.<sup>31</sup>

## 3. Published applications in Corning AFR

### 3.1. Photochemistry

Photochemistry is possible due to an LED system, used from laboratory to industrial scale.<sup>32-34</sup> While each individual wavelength was characterized,<sup>27</sup> the behavior of a multiphase system with photochemistry was also characterized.<sup>35</sup>

#### 3.1.1. Gas photochemistry: Oxygen oxidation.

For alpha-terpinene oxidation, optimizing photochemistry guidelines were published.<sup>36</sup>  $\beta$ -dicarbonyl compounds were enantioselectively oxidized.<sup>37</sup> Sulfured Methionine amino acid was oxidized<sup>38</sup> and the protocol was extended so that mustard gas can be neutralized by air.<sup>39</sup> Use

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of sensitizer on nanoparticles was compared to batch for an oxygen oxidation.<sup>40</sup> Function of terminal N-Methyl groups into aldehydes could also be performed without metals.<sup>41</sup>

### 3.1.2. Materials

Gold nanoparticles can be synthesized, showing the multi-purpose possibility of the reactor.<sup>42</sup> On top of it, daily use of *aqua regia* showed the chemical tolerance of the reactor.

### 3.1.3. Halogen Photo-Chemistry

Iodoperfluoroalkylation of alkenes were carried out.<sup>43</sup> Benzylic bromination reaction was also successfully performed<sup>44</sup>. Another example on G1 scale using NBS was successful.<sup>45</sup>

### 3.1.4. Potentially hazardous species “in situ”

Potentially hazardous species can be generated and used *in situ*, leveraging the inherently safer technology used in continuous manufacturing. Amongst them, Bromine can be generated and reacted *in situ* at laboratory<sup>46</sup> and industrial scales.<sup>47,48</sup> Similarly, nitrosyl chloride can perform photonitrisation.<sup>49</sup> N-Chloroamines were synthesized metal-free by radical addition reactions in continuous flow.<sup>50</sup>

### 3.1.5. Cycloaddition

Selective photoredox transformation can be performed.<sup>51</sup> [2+2] Cycloaddition reaction, supported *in silico*, were performed in G1 reactors.<sup>52</sup> Cerium also catalyzed Cycloalkanols Cycloaddition<sup>53</sup> and functionalize alkanes.<sup>54</sup>

Using renewable source chemicals,  $\gamma$ -butyrolactone were synthesized.<sup>55</sup>

### 3.1.6. Organometallics

Using Nickel as a catalyst, arylhydrazines were synthesized.<sup>56</sup> Using inline NMR monitoring, Nickel Negishi coupling reactions were also carried out.<sup>57</sup> An API intermediate was synthesized this way.<sup>58</sup>

### 3.1.7. Green Chemistry

Direct metal free organocatalytic arylation coupling to aryl bromide was performed.<sup>59</sup>

## 3.2. Thermal Chemistry

### 3.2.1. Classical Chemistry/Batch to Flow

Plant design and economic study of Ibuprofen and artemisinin was evaluated in flow.<sup>60</sup> The use of the appropriate analytical tools (such as Raman spectroscopy) is an asset to ensure a full optimization of process in flow.<sup>61</sup>

Collecting internal data, a Moffat-Swern oxidation was translated from Batch to Flow Chemistry.<sup>62</sup> This showcase highlights the number of possible reactions which can be used in flow. The exothermic chlorination of a compound with thionyl chloride was performed from laboratory to industrial scale both in simulation and experimentally.<sup>63</sup>

### 3.2.2. Synthesis of potentially hazardous chemicals

Using flow reactors, potentially hazardous species can be synthesized using inherently safer technology.

#### 3.2.2.1. Nitric acid use.

Alcohol esterification with nitrous acid, while being an exothermic process, could be carried out successfully in G1 Reactors and turned into synthetically useful alkyl nitrites.<sup>64</sup> Similar

nitration reactions can be performed effectively.<sup>65</sup>

#### 3.2.2.2. Nitrogen/Azide compounds,

While potentially hazardous but synthetically interesting, these reactions have been successfully implemented in AFR technology. Monomethylhydrazine was synthesised.<sup>66</sup> Despite the risks associated with hydrazoic acid, there is a synthesis of Diphenylphosphoryl azide.<sup>67</sup> Using azides, Ritalin was synthesized.<sup>68</sup> Similarly, *in situ*, generated diazomethane was used in a synthetic way.<sup>69</sup> Cyclopropanation was successfully implemented through Design of Experiment strategy.<sup>70</sup>

Benzoic acid alkylation reaction, generating and consuming *in situ* dangerous intermediate species, was performed in flow.<sup>71</sup> Tetrazole coupling reaction was performed, keeping in check all parameters in typically unstable condition.<sup>72</sup>

#### 3.2.2.3. Oxidation

Peracids, unstable species formed in conditions where their stability depends upon a reliable set of unstable conditions, were synthesized effectively.<sup>73</sup> Synthesis of Modafinil was performed smoothly with Hydrogen Peroxide as an oxidant.<sup>74</sup>

#### 3.2.2.4. Use of Gas.

Using oxygen, benzylic oxidation was carried out in metal free and reagent recyclable conditions.<sup>75</sup> Oxygen was also helped with the hydroxylation of ketones and ketamine synthesis.<sup>76</sup>

Ozonolysis, which is potentially hazardous even at a trace level, was performed in a Low Flow

Reactor.<sup>77</sup> A successful case was published at kilo lab scale.<sup>78</sup>

On the other hand, reduction via hydrogenation could be performed, too.<sup>79,80</sup> For a hydrogenation reaction, a system with Pd allowed a temporary Pd deposit *in situ*.<sup>81</sup>

Challenging Bunsen reaction (Gas SO<sub>2</sub>/liquid) requiring precise mixing was industrially implemented.<sup>82,83</sup>

Synthesis of an anti-bacterial agent performic peracid was successfully carried out.<sup>73</sup>

The electrophilic  $\alpha$ -aminohydroxylation of ketones was carried out by preparing *in situ* the 1-chloro-1-nitrosocyclopentane reagent.<sup>84</sup>

### 3.2.3. **Green Process**

Using flow chemistry, a strong emphasis on Green Chemistry is pushed.<sup>85,86</sup>

#### 3.2.3.1. Greener conditions

First, existing applications are optimized in a more ecofriendly way. Tertiary Ketone were hydroxylated without need for metal.<sup>87</sup>

Cyclic organic carbonates were synthesized<sup>88</sup> and solvent-free options were also developed.<sup>89</sup> Solvent free biphasic alcohol oxidation was carried out and scaled up in a Low Flow Reactor.<sup>90</sup> Using bleach, alcohol was oxidized and scale up in a biphasic mixture in a metal free process.<sup>91</sup> LAH reduction of esters into aldehyde was performed in mild conditions.<sup>92</sup> Epoxide nucleophilic opening was used for a coupling en route to the synthesis of an API, telescoping steps and removing DCM as a solvent.<sup>93</sup>

#### 3.2.3.2. Sustainable Material

Synthesis from green glycerol towards oxiranes was performed. Biodiesel could be synthesized from cooking oil.<sup>94</sup> Similarly, biodiesel additive STBE was synthesized from bio-sourced glycerol.<sup>95,96</sup> Total synthesis of Modafinil was entirely performed in flow in 3 steps.<sup>74</sup>

### 3.2.3.3. Biosynthesis

The bioprocess of lipase  $\beta$ -catalyzed isoamyl acetate synthesis was carried out in flow.<sup>97</sup>

### 3.2.4. **Material Chemistry/Nanoparticles**

Iron oxide nanoparticles were synthesized.<sup>98</sup> Further characterization of the equipment and synthesis of iron nanoparticles was successfully carried out.<sup>19</sup> Micro-encapsulation led to smooth, monodisperse and stable components.<sup>99</sup>

Working on asteroids, valuable metals were extracted.<sup>100</sup>

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